

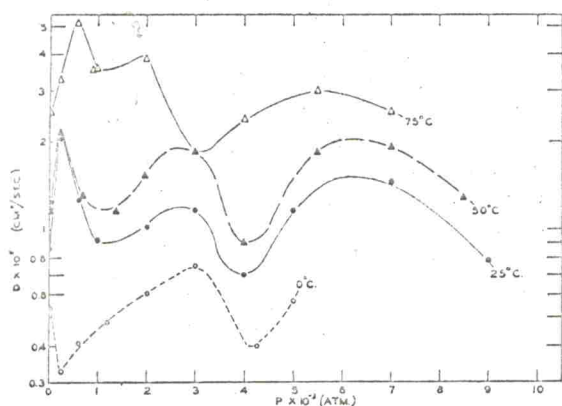


FIG. 1. Diffusion coefficients *versus* pressure, 0.1 N $\text{Hg}(\text{NO}_3)_2$.

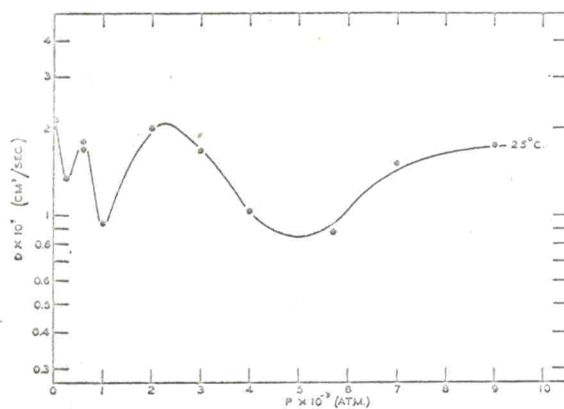


FIG. 4. Diffusion coefficients *versus* pressure, 0.1 N TiNO_3 .

minimum shifts toward lower pressure as the temperature increases. The curves also show a maximum in the pressure range 250–600 atmospheres which moves to

activation entropy from one atmosphere to P are shown in Figs. 7 and 8. Any interpretation of these curves

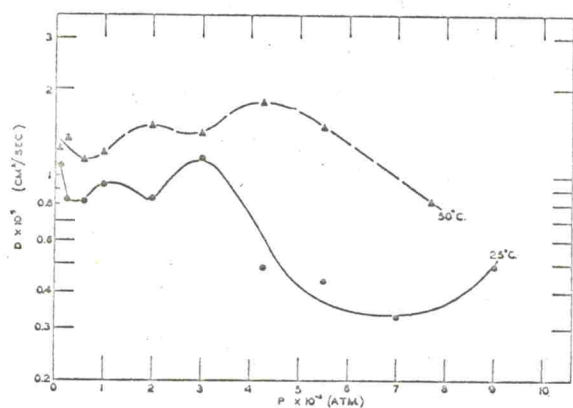


FIG. 2. Diffusion coefficients *versus* pressure, 0.1 N CaCl_2 .

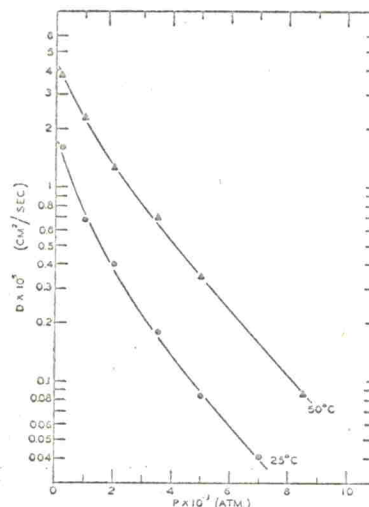


FIG. 5. Diffusion coefficients *versus* pressure, 0.01 M HgCl_2 in *n*-butanol.

lower pressure as temperature decreases and disappears at 0°C.

Curves of activation enthalpy and the increase in

would be dubious because the activation enthalpy and entropy depend on the displacement of two or more

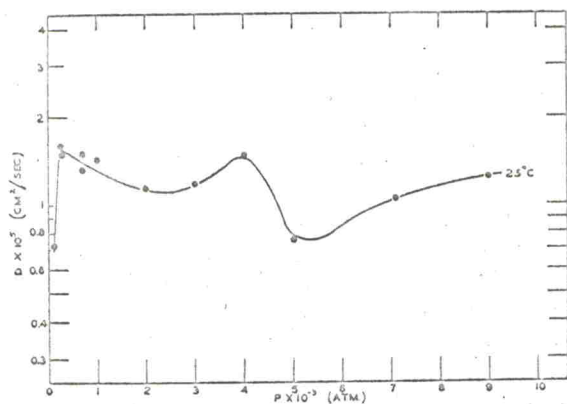


FIG. 3. Diffusion coefficients *versus* pressure, 0.1 N $\text{Ca}(\text{NO}_3)_2$.

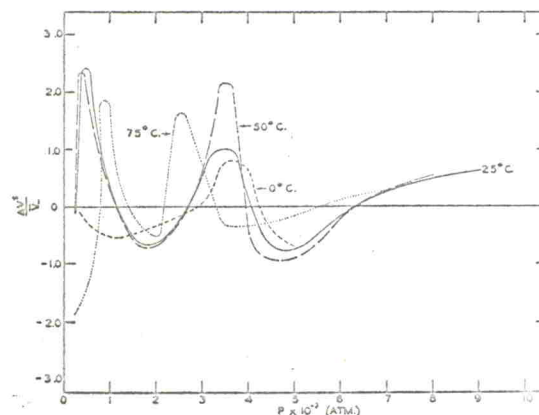


FIG. 6. Activation volume ratios *versus* pressure, 0.1 N $\text{Hg}(\text{NO}_3)_2$.

Observed
 $D \times 10^5$
 cm^2/sec

0.72
1.49
1.60
1.32
1.50
1.43
1.13
1.17
1.47
0.75
1.04
1.24

verage deviation

t and most
tem because
d cover the

0.1 N TiNO_3 ,^a

Observed
 $D \times 10^5$
 cm^2/sec

2.07
1.35
1.80
1.69
0.95
2.00
1.67
1.03
0.88
1.48
1.73

verage deviation

ental work.
from Figs. 1
a coefficient
, and this

DH,^a

Observed
 $D \times 10^5$
 cm^2/sec

1.61
0.68
0.40
0.18
0.084
0.041
3.75
2.29
1.26
0.70
0.35
0.088

age deviation